

Atmos. Meas. Tech. Discuss., referee comment RC1  
<https://doi.org/10.5194/amt-2022-41-RC1>, 2022  
© Author(s) 2022. This work is distributed under  
the Creative Commons Attribution 4.0 License.

## Comment on amt-2022-41

Anonymous Referee #2

---

Referee comment on "Influence of CO<sub>2</sub> adsorption on cylinders and fractionation of CO<sub>2</sub> and air during the preparation of a standard mixture" by Nobuyuki Aoki et al., Atmos. Meas. Tech. Discuss., <https://doi.org/10.5194/amt-2022-41-RC1>, 2022

---

The manuscript describes experimental work to quantify effects that affect the mole fraction of CO<sub>2</sub> in standard air mixtures prepared by gravimetric methods. This certainly is a highly relevant topic within the scope of AMT. The topic has been addressed already by other papers that are adequately referenced. The experimental work has been performed thoroughly, the results of many individual experiments are very consistent and such add valuable quantitative data on the adsorptive loss of CO<sub>2</sub> on high pressure cylinder surfaces and fractionation effects when transferring gas mixtures from one cylinder to another for the purpose of preparing diluted samples. In previous studies of the investigation of such adsorptive effects results it was not always unambiguous if those results might also have been impacted by thermal fractionation. The very good agreement of several decanting experiments performed in this manuscript at different low flow rates yielding similar adsorbed quantities of CO<sub>2</sub> at cylinder walls as others have also published provides further confirmation of the quantitative relevance of this effect.

The results of the second series of experiments on fractionation during gas mixture transfer from one cylinder to an evacuated second cylinder (mother-daughter and dilution experiments) show clearly that the CO<sub>2</sub> mole fraction is altered by the process. It is shown that this effect is quantitatively very reproducible with the experimental setup used for these experiments. This setup normally involves high flow rates for the gas transfer that are probably the cause for the disturbance. Three mother-daughter experiments have been conducted at lower transfer speeds resulting in a lower fractionation. However, this mode of operation has not been further employed at the panel utilized apparently is not made for such low flows. At this point the description of the experimental setup has not been sufficiently detailed to understand this. Isotope ratio mass spectrometry has been used to study the relationships in fractionation of various isotope or molecule pairs (e.g. <sup>29</sup>N<sub>2</sub>/<sup>28</sup>N<sub>2</sub>; <sup>34</sup>O<sub>2</sub>/<sup>32</sup>O<sub>2</sub>; Ar/N<sub>2</sub>; O<sub>2</sub>/N<sub>2</sub>;..) and compare these measured relationships with literature results as an indicator for the kind of diffusive process causing the fractionation. These measurements apparently are not trivial and neither could a convincing conclusion be drawn from these data nor appears the discussion of these results very substantial.

There is a problem with the language being not sufficiently clear in several sentences throughout the manuscript that requires a thorough revision. For some key sentences I

have provided suggestions below but this is certainly not exhaustive.

Concrete points that I feel the authors need to further elaborate or correct before final publication are listed here below.

lines 40-44: The conclusion drawn here is misleading and therefore should not remain. The quoted WMO compatibility goal is part of WMO recommendations that request to refer to the WMO CO<sub>2</sub> mole fraction scale. By using this approach the absolute accuracy of the scale is of less importance and not a major contribution to the limit of the WMO GAW network compatibility. However, the two comparison studies that are referred to (Tsuboi et al, 2017 and Flores et al 2019) both comprise participating laboratories that mostly report on different scales. In the Japanese study it appears that for each lab the consistency of offsets of three comparison samples relative to NIES is very good for all the labs and the agreement between AIST and TU (both reporting on the TU2010 scale) is within the WMO compatibility goal. The comparison results of the CCQM Key Comparison (Flores et al 2019) provides the result of measurements performed at the BIPM of standard gases provided by the participating NMIs and the WMO Central Calibration Laboratory. Both references therefore do not provide any information on the compatibility of measurements by different laboratories relating to the WMO scale that might limit an evaluation of the carbon cycle.

l.81: please insert reference: ".. into another cylinder (Hall et al. 2019)"

p.5 l. 120: " some of these mixtures were purchased from a gas supplier (Japan Fine Products, Japan), while others were prepared at our laboratory.". This does not state if the gas mixtures are synthetic (which becomes clear later in the manuscript) or based on real ambient air. This should be stated explicitly here. What is confusing is the statement made at a later point (p. 6 l. 143) that "Atmospheric values of N<sub>2</sub>, O<sub>2</sub>, Ar ... were used as values of N<sub>2</sub>, O<sub>2</sub>, Ar.". This does not appear valid for synthetic mixtures.

p.5/6 l. 127f: I guess that I do not understand the following sentence and it is not in clear agreement with Fig 1a: "The transfer speed was controlled using the only diaphragm valve with the daughter cylinders calculated roughly from the transfer time and volume."

It appears that the full pressure of the mother cylinder is applied to the system with the evacuated daughter cylinder at the other end. This should cause a dramatically large flow entailing big temperature changes in the Mother and Daughter cylinders. There are several valves depicted in the figure that the gas is passing through. I do not know the type of diaphragm valve so what restriction it does provide. The one pressure gauge in Fig. 1a is downstream of diaphragm valves from the Mother cylinders and upstream of the diaphragm valve of the Daughter cylinders so it is not clear to me how it is checked when the transfer volume has been reached. How is the transfer time then measured? How long is the transfer typically taking? In Fig. 1a there are two cylinders named Mother and one Daughter. Please revise to make it clear.

p 6. l 141: It would be helpful if the equation 1 could be explicitly derived.

p 6. l. 149: Could you please specify what is G1-grade air?

p.7 l. 158: Was there any significant humidity detectable in the air?

p.8 l. 184ff: It is not explained why the 0.8L cylinder is moved from V6 to V10. Is it in order to avoid any CO<sub>2</sub> getting lost in the the valves V6-V9? If so, how has been assured that no additional CO<sub>2</sub> that might have adsorbed in valves V1 - V7 is later flushed into the 10 L cylinder during the 300 pressure expansion cycles with the dilution air?

p.8 l. 197: Was this preparation done using the set-up displayed in Fig. 1a?

p.9 l. 215f: "The output of the Picarro G2301 was calibrated using standard mixtures prepared by the one-step dilution." In section 2.2.3 it is not specified how many one-step dilution standards are prepared in parallel. This information is provided at a later stage but number of calibration points and the CO<sub>2</sub> mole fraction range covered should also be mentioned here.

p.9 l. 223: "However, diffusive fractionation in the transfer of **the** source gas is likely to have a larger impact on the CO<sub>2</sub> molar fractions than the bias by the **sorption processes**." This appears like an anticipation of the results the study is concluding with. It also relates to the large flow rates employed in this experimental set-up.

p.11 l. 250: Omit the first sentence ("The standard deviation is..")

p.12 l. 297-299: This sentence needs to be corrected as follows: "**Lighter molecules preferentially escape from the orifice relative to heavier ones** because the rate of effusion is inversely proportional to the square root of the mass".

p.13 l. 306: Is the diffusive process quick enough to cause a fractionation because of the pressure gradient between Mother and Daughter cylinder entail very high bulk flow rates?

p.13 l. 307f: I would think this sentence: "Additionally,.. " should be omitted, as all these

diffusive processes are based on molecular mass differences.

p.13 l. 313f / Fig. 4 / Fig. 1b: The description of the panel used for the mother-daughter experiments depicted in Fig. 1b does not make clear how the transfer speed is controlled but Fig. 4 clearly indicates that in the experimental series transfer speed and transfer volume were independently controlled. Therefore, please do clarify in Fig. 1a or section 2.1 by what means (which type of valves) the flow can be regulated.

p.14 l. 333: The statement "...but it becomes significantly weaker as the transfer speed decreases" suggests a continuous relationship which the data do not show. Rather change to "...but it becomes significantly smaller below flow rates of 19 L/min"

p.14 l. 336: Why would you consider Knudsen diffusion to play a role? The orifice of the valves certainly will exceed the free path length of molecular collisions. Omit "effusion".

p.14 l. 344: This sentence poses again the question: how has the flow been regulated to perform the experiments resulting in the transfer speeds  $< 3\text{L/min}$  (last three data points of Table 1 and open triangles in Fig. 4) ?

p.15 l. 333: Does the statement "...using the values in Table 1." include alle the values (that is what is written) or only values with transfer speeds  $\geq 19\text{ L/min}$  (that would make more sense to me, as apparently the fractionation factor is not constant below some threshold between 3 and 19 L/min)?

p.15 l. 372: As stated above, Knudsen diffusion ("effusion") will not be of relevance in the mother-daughter experiments. Therefore, I find it rather confusing than helpful to include it as another dotted line in the Fig. 5 plots.

p.15/16 l. 365 - 394: The mass spectrometric analysis of isotope or molecule pair ratios appears to be a powerful tool to ascertain the fractionation process. Yet, as the presentation and discussion of the results currently are and with the conclusion drawn the reader feels left behind without much of a clue if the data confirm any assumption or not.

p.15/16 l. 376 - 377: It is surprising that only for the two cases where the relationships for thermal diffusion from Langenfelds and Ishidoya investigation are in agreement the latter are not included in Fig. 5. This solid red line should be either added or it should be explained why it is not. The slopes of the other delta relationships in Figure 5 are varying quite a bit between the two references. The authors attribute this to "some factors that were not considered in the theoretical model " (= dotted line from Langenfelds 2005). However, the spread of the measured data presented here is of similar magnitude than

this disagreement. This might therefore also point to an uncertainty in the accuracy of the experimentally determined literature value. Unfortunately, this discrepancy is particularly large for the  $D_{CO_2/N_2}$  vs  $D_{^{29}N_2/^{28}N_2}$  and the authors acknowledge: "Here, note that the deviation of the experimental thermal diffusion for  $\delta(CO_2/N_2)$  has large uncertainty and requires further experiments.". It would help the reader if that statement could be either more quantitative or be explained more in detail. Is the underlying cause for that stated uncertainty an issue that might also apply to the measurements made in this study? Could information be provided on the measurement precision for the displayed delta-values (error bars on the black dots).

p.15 l. 380: The following statement does not convince me: "The fact that the deviations of  $\delta(CO_2/N_2)$  are close to the experimental thermal diffusion, indicates that the fractionations occurred by thermal diffusion." The agreement of the delta relationship is closer to the value from the Ishidoya studies than to the quoted Langenfelds number. However, on the one hand, their had just been doubts raised in l. 374 on that Ishidoya number, and on the other hand the observed deviation of the  $\delta_{CO_2/N_2}$  ratio is also clearly larger than in the Ishidoya et al study.

In my opinion it does not help to vaguely point to "..unknown fractionation mechanisms.." or refer to diffusive processes occurring in firm that do not seem of any relevance to processes occurring in the mother daughter experiment with enormous bulk flow rates and without very small orifices.

p.18 l. 437: The following sentence is incomprehensible: " These contributions are negligible to the increase because all cylinders act similarly.. ". What does it refer to (which contributions)?

".., although the fractionation in the transfer of the 2nd gas mixture into the daughter cylinder and ... also affect  $CO_2$  molar fraction in the 3rd gas mixture." Why *although*? This is what has been explained to be the goal: to backtrack the impact of fractionation on the 2nd mixture by analysing the 3rd mixture.

I do not see the need for the sentence, but if I miss the message it should be reformulated.

p.18 l. 445: The following sentence is confusing "The deviations increased by  $0.25 \pm 0.10 \mu\text{mol mol}^{-1}$ ..." in that respect, that  $CO_2$  increased but not the absolute value of the offset from the gravimetric value: at 11.5 MPa there is a deviation whereas at 2 MPa there is no deviation. You could either set the deviation from the gravimetric value of the 3rd mixture produced at 11.5 MPa pressure as reference ( $y\text{-axis} = 0$ ), or re-phrase the sentence as e.g. "The known negative offset from the gravimetric value caused by the fractionation process in the gas transfer during the 3rd gas mixture production is observed for the 3rd prepared from the 2nd at 11.5 MPa. With decreasing 2nd pressure to 1.1 MPa  $CO_2$  increased in the 3rd gas mixture by  $0.25 \pm 0.10 \mu\text{mol mol}^{-1}$ "

p.18 l. 447ff: I am not sure if I understand what has been made here : "...we estimated the fractionation factor..": have the six data points shown in Fig. 8a and Eq. 4 been used to derive a new fractionation factor by this? Should be re-phrased.

p.20 l. 482-482: Why is this described here again? Is this different from the calculation of measured values made before?

p.20 l. 490-492: I would remove the two concluding sentences here. They are well placed in the next section.

p.21 l. 504: The data in Fig. 4c do **not** show that the decrease in CO<sub>2</sub> is "weakened significantly with decreasing of the transfer speed". In this data set the decrease is independent from the flow rate in the range from 19 - 216 L/min! Needs to be re-phrased, e.g. "The decrease in CO<sub>2</sub> molar fractions in the daughter cylinders does not depend on the transfer volume and initial pressure. This fractionation effect is neither depending on the transfer speed at flow rates exceeding 19 L/min but significantly reduced at lower flow rates ...

p.21 l. 513f: I suggest to rephrase: "The fractionation caused the CO<sub>2</sub> molar fraction to increase and decrease. The reproducibility of CO<sub>2</sub> molar fractions in gravimetric standard mixtures will suffer as a result." to "Fractionation at different stages of a multi-step dilution can result in CO<sub>2</sub> increases as well as in CO<sub>2</sub> decreases of the final gas mixture. This affects the reproducibility and accuracy of CO<sub>2</sub> molar fractions in gravimetric standard gases"

p.32 Fig. 5 top panel, y-axis: CO<sub>2</sub>/<sup>28</sup>N<sub>2</sub>)(permeg)

## Language

p.1 l. 16.: ..only once. **It** indicates...

p.1 l.25: Needs to be re-phrased to become clear, as the following wording is probably wrong: ""The fact that the *CO<sub>2</sub> molar fraction weakened* significantly...". Is the intention to state that the CO<sub>2</sub> mole fraction (in the escaping?) gas was reduced when the transfer speed was reduced? I presume it means to say: "The fact that the CO<sub>2</sub> fractionation effect was less significant as the transfer speed decreased.."

p.1 l.27: The following would be easier to comprehend: "Experiments were conducted where a CO<sub>2</sub> in air mixture was emitted from a cylinder to evaluate the CO<sub>2</sub> adsorption..."

p.2 l.29: rephrase, e.g. "...the CO<sub>2</sub> molar fraction in the exiting gas mixture increased by  $0.16 \pm 0.04 \mu\text{mol mol}^{-1}$ ."

p.2 l. 39: It would be appropriate to provide references for the primary standard laboratories

P.2 l45: ..that CO<sub>2</sub> adsorbed **on** the internal...

p.3 l. 65: Langen**felds**...

p.3 l.76f: Rephrasing the sentence would improve its clarity, e.g. : "...daughter cylinder causing an increase in the CO<sub>2</sub> molar fraction in the remaining source gas in the mother cylinder".

p.3 l. 77: Rephrasing the sentence would improve its clarity, e.g. : "This could be a factor that deteriorates the reproducibility of the assigned CO<sub>2</sub> molar fractions because CO<sub>2</sub> molar fractions in the prepared standard mixtures are biased by the decrease and increase in CO<sub>2</sub> in the transferred gas mixture and the remaining pre-mixture, respectively."

p.4 l. 88-90: The need of the sentence is not clear to me. I do not understand what the "**Although**.." aims to qualify.

p.4 l. 99: ".. estimate CO<sub>2</sub> adsorption **on** the internal..."

p.5 l. 13-15: "The Picarro G2301 output was linearly calibrated using one standard mixture containing atmospheric CO<sub>2</sub> levels with a standard uncertainty of less than  $0.1 \mu\text{mol mol}^{-1}$ ."

p.12 l. 280: "The fractionation of CO<sub>2</sub> and air **during** the transfer of a gas mixture with atmospheric CO<sub>2</sub> level has can be caused not only **by** the diffusive process but also the adsorption process."

p.12 l. 289: "Therefore, the fractionation of CO<sub>2</sub> and air **is expected to result from** the diffusive fractionation process based on the three types of diffusion,.."

13 l. 322: change "However,.. " to "Correspondingly,.."

14 l. 340: ".. of standard mixtures with **accurate** atmospheric CO<sub>2</sub>..."

p.21 l. 510: "We demonstrated that CO<sub>2</sub> molar fractions in standard mixtures by three-step dilutions decreased by  $-0.207 \pm 0.060 \mu\text{mol mol}^{-1}$  from gravimetric values based on source gas fractionation, which is greater than the compatibility goal of  $0.1 \mu\text{mol mol}^{-1}$ ."